



CLASS :VIII

Subject:CHEMISTRY
Topic:MATTER

Date:29/04/2020
Time:35 mins

Worksheet No.:2

Good morning children in the last class we have seen the various types of matter .

Today I am going to tell you the various kinetic theory of matter and how matter is different from each other based on these postulates

KINETIC THEORY OF MATTER

*It states that the matter is composed of large number of small particles,which are in constant motion.

POSTULATES OF KINETIC THEORY OF MATTER

*Matter consists of minute particles which may be atoms ,ions,or molecules.

*The particles are attracted to each other by a force called **INTER MOLECULAR FORCE OF ATTRACTION**.

*The shape of matter can be differentiated on the based on the **INTERMOLECULAR SPACE**.

*All particles are in constant random motion which is directly proportional to kinetic energy .

*A change in phase occurs when the energy of the particle is changed.

DIFFERENCE BETWEEN SOLID/LIQUID/GAS

SOLID	LIQUID	GAS
1.Molecules are closely Packed .intermolecular Force is stronger.	1.Molecules are little apart from each other the Intermolecular force is less Than solid.	1.Molecules are far apart from each other. Intermolecular Force is very weak.
2.Intermolecular space Is minimum.	2.Intermolecular space is More than solids.	2.Intermolecular Space is maximum.
3.They can vibrate about Their mean position. Kinetic energy is Minimum.	3.The molecules move from One position to the other . Kinetic energy is little more Than solids.	3.The molecules keep moving randomly. Kinetic energy is Maximum.
4.Small expansion on Heating.	4.expands little more than Solids.	4.shows maximum Expansion.

[Copy the questions and solve them on a sheet of paper date wise. Keep the worksheets ready in a file to be submitted on the opening day.]

1.Define :

intermolecular space and intermolecular force of attraction.

2.State five postulates of kinetic molecular theory.

3.Differentiate between solid,liquid,gas.[based on compressibility,intermolecular space,kinetic energy,expansion.]
